Welcome to another experiment.

Today I am going to demonstrate a precipitation reaction.

I am going to perform the reaction between **lead nitrate Pb**(**NO3)2** and **potassium iodide** both of them are white colored salts and we also require 3 test tube and some water.

So let’s start the experiment.

First of all I am going to take some **lead nitrate** in a test tube and in another test tube I am going to add some **potassium iodide.** Now in this **potassium iodide** I am going to add some **water** and shake it.

I am also going to add some water In **lead nitrate** and shake it well. Now I am going to mix **lead nitrate** Liquid with **potassium iodide** liquid. Here we can see that the precipitates have formed and it is yellow in color. Now Let’s see the reaction,

**KI + Pb(NO3)2  -> PbI2 +KNO3**

Here the compounds of potassium dissolves in water. That means the substances that does not dissolve in water which is **PbI2** forms a precipitate. So **PbI2** precipitates which are yellow in color. And I should let it settle down for at least an hour so that we may see the precipitates.

Thank you Everyone.